

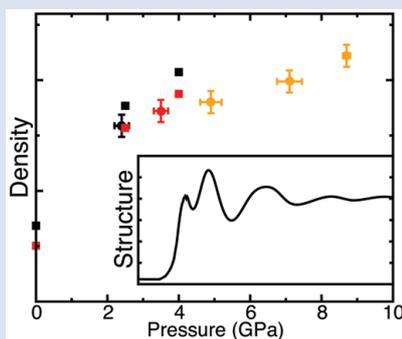
Properties of molten CaCO₃ at high pressure

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Abstract



We report here the structure of molten CaCO₃ studied by *in situ* X-ray diffraction using a Paris-Edinburgh press up to 8.7 GPa. Variations are observed in the medium range order with shrinkage of the intermolecular contributions, reflecting higher packing efficiency of the carbonate molecules. Density of the melt is obtained from the radial distribution functions, assuming a constant coordination number of 3 for the C-O contribution. Bulk modulus values increase by a factor of 2 over the experimental pressure-temperature range, reaching a value similar to that of underlying crystalline phases at the highest pressure investigated. These are the first direct density measurements of compressed CaCO₃ melt; they agree well with recent *ab initio* predictions (Li *et al.*, 2017), which implies that the reported flattening of the melting curve followed by a slightly negative curve above 8 GPa cannot be due to crystal-melt density inversion. Instead, the enthalpy of fusion is likely responsible, and we note the peculiar high diffusivity of oxygen in crystalline calcite-V in this regard, with further implications for the mobility of carbonate melts at depth.

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Introduction

The geological importance of carbonate melts is linked to their low melting points compared to silicates, and hence their key role in triggering melting at depth (Jones *et al.*, 2013). As such, the properties of carbonate melts are key in processes ranging from incipient melting at the oceanic lithosphere-asthenosphere boundary and lubrication of tectonic plates (Gaillard *et al.*, 2008), to pervasive metasomatism of the continental lithospheric roots (Foley, 2008).

Experimental measurements of the physical properties of molten carbonates under high pressures (P) are very scarce however. This contrasts with the advances in knowledge on silicate melts under P over the last decade, as reviewed in Kono and Sanloup (2018), and is due to the extreme difficulty in confining such low-viscosity melts for long durations. Structural data at ambient P have been collected on Li₂CO₃, K₂CO₃, LiKCO₃ melts using neutron diffraction (Kohara *et al.*, 1998), and Na₂CO₃ using X-ray diffraction (Wilding *et al.*, 2016). The effect of P on melt compaction, *i.e.* the bulk modulus, has been assessed for alkali carbonate melts by combining ambient P density and acoustic velocity measurements (Liu and Lange, 2003), and from the melting curve for K₂CO₃ (Liu *et al.*, 2007). *In situ* high P density measurements are restricted to two P points for molten K₂Ca(CO₃)₂ (2.75 g·cm⁻³ at 2.5 GPa-950 °C, 2.58 g·cm⁻³ at 2.5 GPa-1150 °C, and 2.80 g·cm⁻³ at 4 GPa-1050 °C) using the falling sphere technique (Dobson *et al.*, 1996). These first *in situ* density measurements were

complemented by a study of the viscosity of K₂Ca(CO₃)₂ and K₂Mg(CO₃)₂ melts up to 5.5 GPa (Dobson *et al.*, 1996). Such viscosity measurements were very challenging at the time, *i.e.* before the advent of high speed cameras required to accurately capture rapidly falling spheres. The ultralow viscosity, *i.e.* two orders of magnitude lower than for molten basalt at 1900 K, was nonetheless confirmed for calcite and dolomite melts (Kono *et al.*, 2014), with no P effect observed up to 6.2 GPa. This latter study also reported the structure of CaCO₃ melts but information was restricted to distances greater than 2 Å, excluding the C-O bond.

In contrast, there are several reported melting curve measurements, as these do not require long experiments and sample integrity can be preserved. Of particular interest is the report of a flattening of the CaCO₃ melting curve at 8 GPa followed by a slightly negative slope up to the calcite-aragonite-melt triple point at 13 GPa (Li *et al.*, 2017). This evolution is unique amongst carbonates, with stronger and positive slopes reported for Na₂CO₃ and MgCO₃ (Li *et al.*, 2017). The flattening of the CaCO₃ melting curve was attributed to a density crossover between solid and liquid CaCO₃. However, such crossover is only predicted theoretically near 17 GPa, and this discrepancy was attributed to approximations in the *ab initio* calculations, highlighting the need for further experimental measurements of melt density.

We report here structural and density data on molten CaCO₃ from *in situ* X-ray diffraction up to 8.7 GPa and 2073 K, an experimental P - T range that corresponds to upper mantle

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conditions. Attempts to collect data on molten Na_2CO_3 , K_2CO_3 , and carbonate mixtures failed, and consequently no results are reported here for these compositions. This might indicate that these compositions have even lower viscosities than pure CaCO_3 melt, resulting in escape of the molten carbonates from the graphite capsule.

Structural Data

In situ energy-dispersive X-ray diffraction data were collected under high P - T conditions (Table S-1) using a Paris-Edinburgh press. The structure factor, $S(q)$, was derived from the X-ray diffraction patterns (Supplementary Information, SI). The $g(r)$ radial distribution function, a measure of the probability of finding an atom as a function of the radial distance r , was obtained by Fourier Transform of the spline smoothed $S(q)$.

$$g(r) - 1 = \frac{1}{2\pi^2 n_0} \int_0^\infty q [S(q) - 1] \sin(qr) dq, \quad \text{Eq.1}$$

where $n_0 = \frac{\rho N_A}{M}$ is the atomic density in atoms per \AA^3 , N_A is

Avogadro's constant, $M = 100.1$ g/mol is the mean atomic molar mass of calcite, and ρ its mass density (see below). The first peak in $g(r)$ is the C-O intramolecular contribution, which

is not expected to change significantly over this P - T range. The second peak in $g(r)$ corresponds to the sum of the O-O and Ca-O first neighbour contributions at ~ 2.3 \AA , and the third peak to the sum of the second Ca-O and first Ca-Ca contributions (~ 4.0 \AA). The most pronounced changes in $g(r)$ (Fig. 1b) with P are at 3 \AA , with an increase of intensity related to the intermolecular O-O distance, and the growth of a shoulder around 3.45-3.5 \AA related to the Ca-Ca intermolecular contribution. These observations are consistent with the shrinkage of interatomic distance, most pronounced for Ca-Ca (from 4 \AA at ambient P), and increase in intensity predicted for both contributions by molecular dynamics (MD) calculations (Vuilleumier *et al.*, 2014), as a result of intermolecular packing upon compression. In contrast, intramolecular distances do not change significantly. The intermolecular packing also explains the observed shift of the first sharp diffraction peak in $S(q)$ (Fig. 1a). Concomitantly, MD simulations indicate a gradual increase of the number of carbonate ions around Ca from 6 as in calcite to 9 as in aragonite. It is not possible to estimate the Ca-C coordination number from the present data due to the prominence of the major contributions of neighbouring Ca-O and Ca-Ca. However, correspondence between Vuilleumier *et al.* (2014) MD simulations and our results supports their conclusions, *i.e.* the P -evolution of the liquid structure smoothes out the abrupt changes occurring across the underlying crystalline calcite V-aragonite transition.

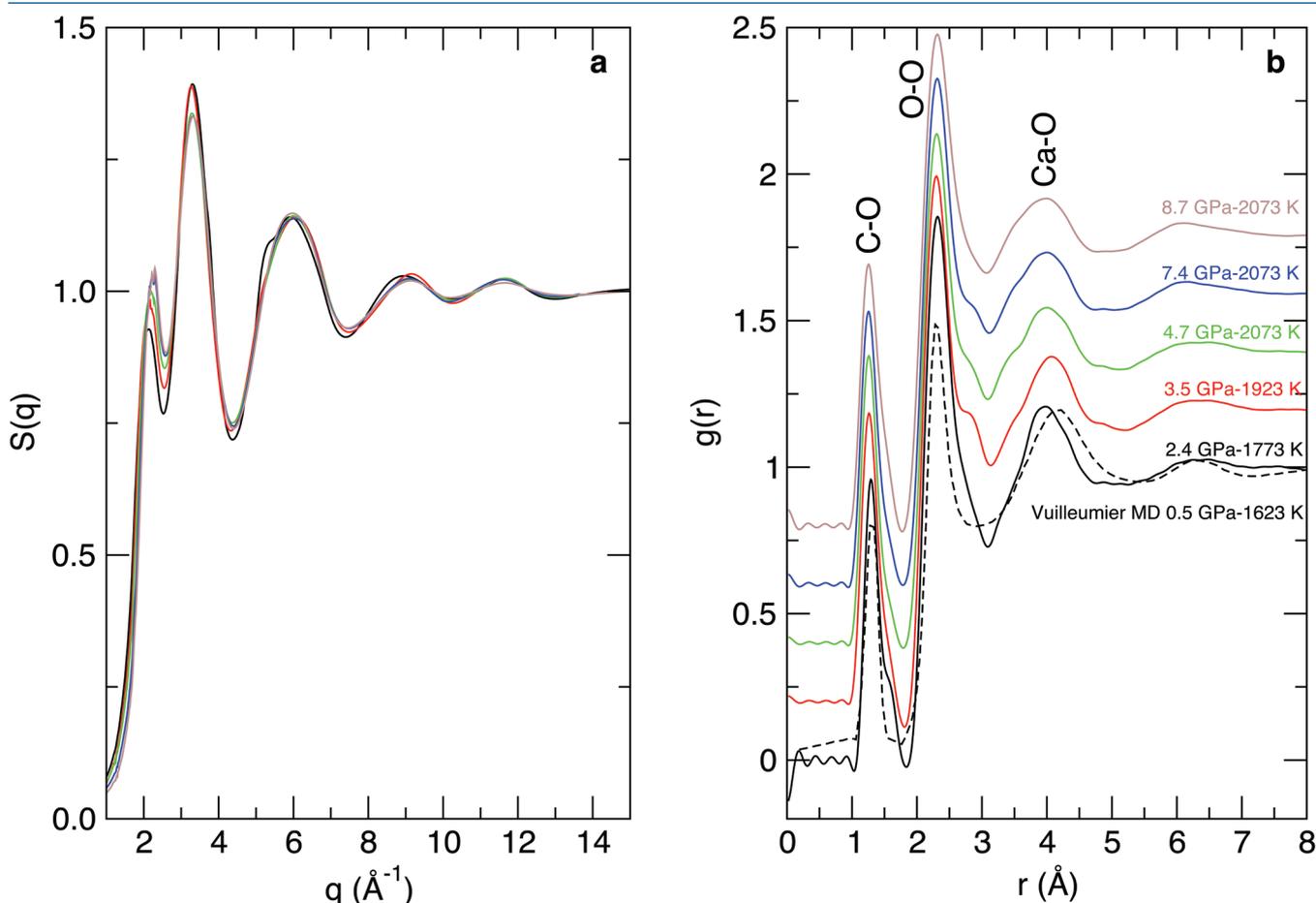


Figure 1 (a) Structure factors, $S(q)$, of molten CaCO_3 ; curves are stacked to see better the evolution with increased P - T conditions (given on the right panel); the main change affecting $S(q)$ (Fig. 1a) is the shift of the first sharp diffraction peak (FSDP) towards higher reciprocal distances, up to 2.28 \AA^{-1} at 8.7 GPa which corresponds in the real space to a characteristic mid-range order distance, $2\pi/q_{\text{FSDP}}$, of 2.76 \AA . (b) Corresponding radial distribution functions (plain curves), $g(r)$, compared to MD simulations (dashed curve; Vuilleumier *et al.*, 2014).

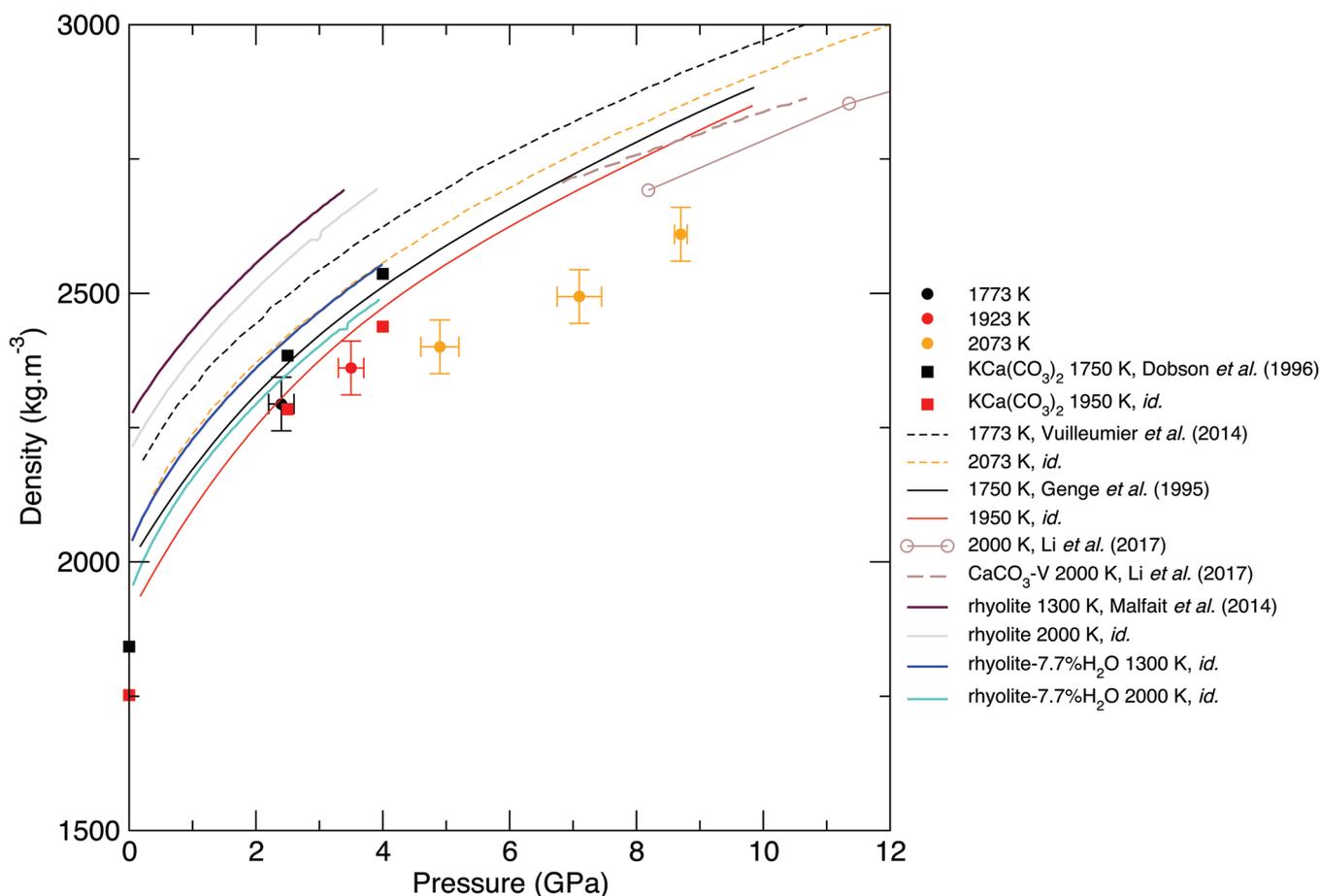


Figure 2 Density of molten CaCO₃ from experiments (circles: this work, squares: Dobson *et al.*, 1996) and theoretical calculations (Genge *et al.*, 1995; Vuilleumier *et al.*, 2014; Li *et al.*, 2017), compared to the seismological PREM model (Dziewonski and Anderson, 1981), crystalline calcite V (Li *et al.*, 2017), molten hydrous and dry rhyolite (Malfait *et al.*, 2014), molten hydrous and dry basalt (Sakamaki *et al.*, 2006).

Density Evolution

The density of CaCO₃ melts (Fig. 2) is calculated from the area below the C-O contribution to the radial distribution function (Eq. 1 and SI). C is coordinated by three O in the studied *P-T* range, but C becomes tetrahedrally coordinated above 50 GPa in crystalline MgCO₃ (Oganov *et al.*, 2008) and above 76 GPa in crystalline CaCO₃ (Pickard and Needs, 2015). How this 3-fold to 4-fold transition translates at very high *P* in the molten state remains to be investigated. The calculated density values are consistent with MD simulations by Genge *et al.* (1995), although the mismatch with more recent simulations (Vuilleumier *et al.*, 2014) is larger, but seems to be in agreement with *ab initio* calculations from Li *et al.* (2017).

It is difficult to determine the isothermal bulk modulus, K_T , by fitting an equation of state to only three data points on the 2073 K isotherm. An alternative method to assess K_T (Egelstaff, 1994) is by extrapolating $S(q)$ to $q = 0 \text{ \AA}^{-1}$ (Fig. S-2

and SI) as $\lim_{q \rightarrow 0} S(q) = \frac{nTk_B}{K_T}$, where k_B is Boltzmann's constant.

This method has been previously benchmarked for molten fayalite using the same experimental set up (Sanloup *et al.*, 2013). Resulting K_T values are provided on Figure 3, along with a correction of the two lowest *P* points to the 2073 K isotherm using $dK_T/dT = 0.02 \text{ K}^{-1}$ from Genge *et al.* (1995). At 8.7 GPa, the bulk modulus ($80.6 \pm 3.5 \text{ GPa}$) is identical to that of crystalline aragonite (79.5–81.3 GPa) using $K_{T,0}$, K' and dK_T/dT values provided in Litasov *et al.* (2017), and although no high *T* data

are available for calcite V, $K_{T,0}$, K' for aragonite is similar to that of calcite I (Redfern and Angel, 1999). Compared to silicate melts (Fig. 2), molten CaCO₃ is less compressible than basalt (Sakamaki *et al.*, 2006) especially at high *P*, but is only approximately 5 % less dense than hydrous rhyolite (Malfait *et al.*, 2014), with a similar *P*-evolution (although the latter has only been investigated up to 4 GPa).

Discussion

Our results confirm the liquid density of CaCO₃ estimated from *ab initio* calculations (Li *et al.*, 2017). There is no experimental data on the thermal equation of state of calcite V due to its very limited *T*-stability field (<200 K) and strong recrystallisation processes at high *T* (Suito *et al.*, 2001), but the good agreement between experimental and theoretical estimates of liquid density lends support to theoretical calcite V equation of state (Li *et al.*, 2017). Density crossover between solid and liquid phases may hence only occur at 17 GPa as predicted (Li *et al.*, 2017). The reason for the flattening of melting curve around 8 GPa, followed by a slightly negative slope up to the calcite V-aragonite-liquid triple point, may consequently be related to changes in the fusion enthalpy rather than a density crossover. Calcite V has a very peculiar molecular structure, with the oxygen sublattice being molten and the oxygen atoms diffusing along an undulated circular orbit (Ishizawa *et al.*, 2013). Such high diffusion of oxygen in crystalline CaCO₃ phases is unique to calcite, and in particular, it does not occur

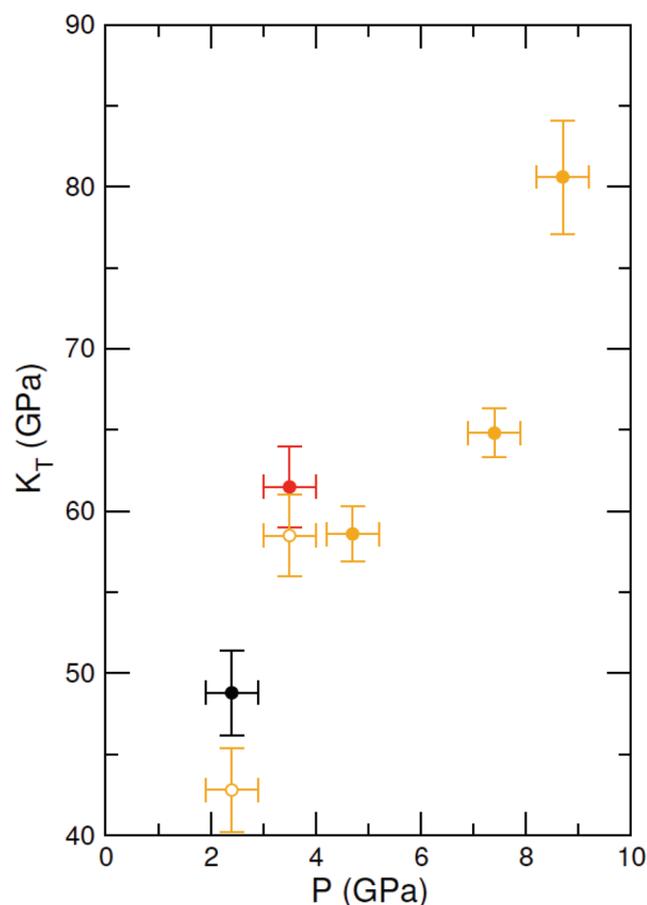


Figure 3 Bulk modulus of molten CaCO_3 as a function of P (black: 1773 K, red: 1923 K, orange: 2073 K, orange empty circles: 1773 K and 1923 K data corrected for $T = 2073$ K).

in aragonite. Although it is not possible to probe this by X-ray diffraction, it is expected that the change of oxygen mobility also occurs in the molten state from calcite-like at low P to aragonite-like at high P . It appears this does not affect the viscosity of the melt *vs.* P (Kono *et al.*, 2014), as the latter is controlled by the diffusivity of the cations. Due to the gradual nature of the structural changes in the molten state, this highly diffusive state will decrease in the melt at lower P than it does in the crystalline state. Enhanced diffusive motion of oxygen in calcite V is consequently expected to cause a large enthalpy decrease upon melting into an aragonite-like high P liquid. Similarly, a strong increase of the melting enthalpy ΔH (*i.e.* $T\Delta S$) of water has been measured above 15 GPa and up to the ice VII-superionic ice X transition, associated with enhanced diffusion of H atoms in the liquid (Goncharov *et al.*, 2009).

The unique feature of the melting curve of CaCO_3 amongst carbonates therefore seems related to the particularity of calcite V. The existence of a P range with negative melting curve related to variation in fusion enthalpy implies that adding heat to the system will freeze it. Corresponding geodynamical settings could be slab-derived melts entrained in upward plumes rising from the transition zone, as proposed for the generation of deep diamonds (Bulanova *et al.*, 2010). Deep carbonate melts cannot indeed derive from melting in the peridotite + CO_2 system but from subducted carbonaceous sediments or from metasomatised eclogites (Hammouda and Keshav, 2015), keeping in mind that carbonatitic melts produced above 15 GPa get Mg-enriched while Ca-rich melts ($\text{Ca}\# = 0.68\text{--}0.70$) are reported in the 8–14 GPa range (Thomson *et al.*, 2016), coincidental to the negative CaCO_3 melting slope. The nature of deep diamond inclusions also

points to the presence of Ca-rich protoliths or calcic melts at depths exceeding 300–350 km (Korsakov and Hermann, 2006; Bulanova *et al.*, 2010). Freezing of melts entrained in a plume might then explain the complex, broken and heavily deformed shapes of deep diamonds, indicating formation in a very viscous environment.

Density and structural data are yet to be collected on other carbonate melts, including alkali compositions, in order to assess the properties of all natural compositions. This will require using faster density/structural probes, due to the challenge of confining such inviscid and highly reactive melts under pressure.

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Additional Information

Supplementary Information accompanies this letter at <http://www.geochemicalperspectivesletters.org/article1813>.



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